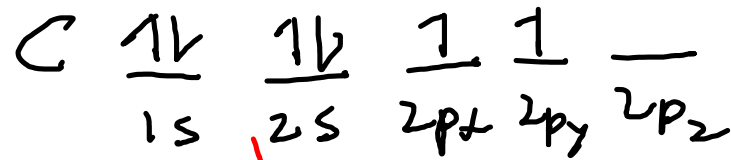
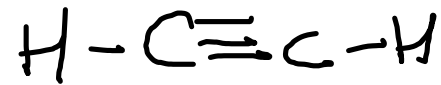
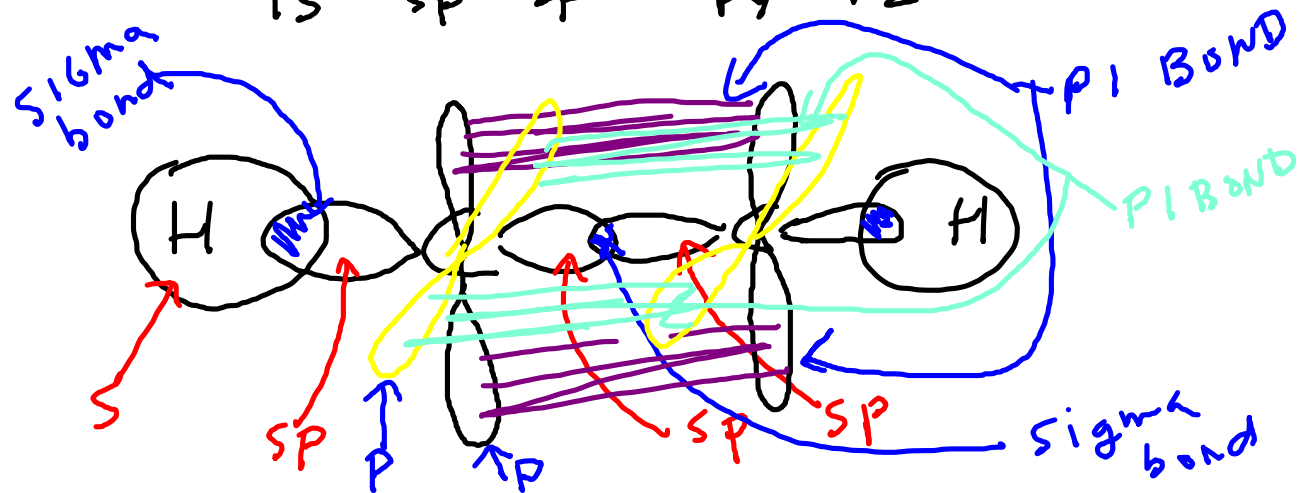
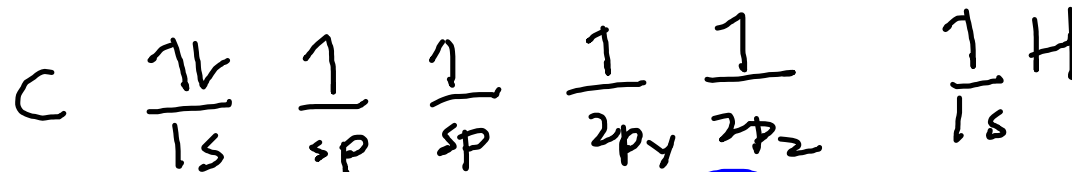


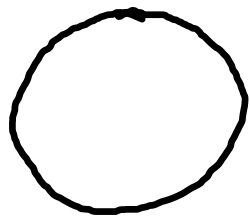
MOLECULAR ORBITALS



HYBRIDIZE



SHAPES OF ORBITALS



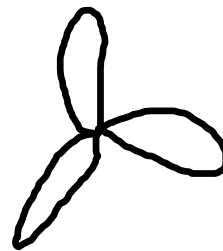
s



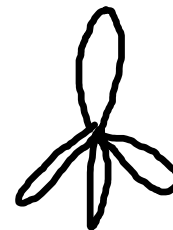
p



sp



sp²

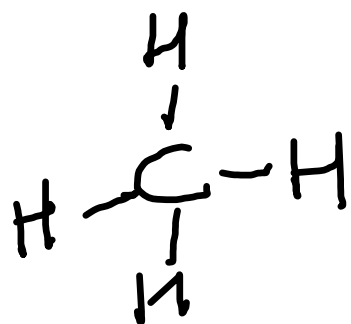


sp³

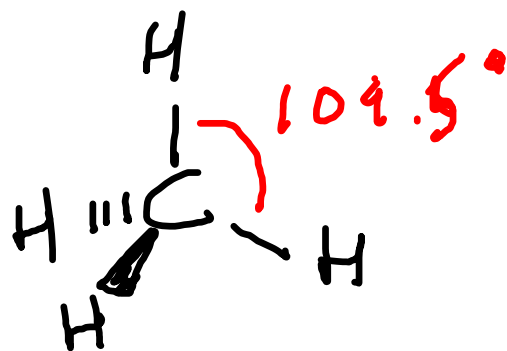
HYBRID	GROUPS OF e ⁻	ANGLES	SHAPE
SP	2	180°	LINEAR
SP ²	3	120°	TRIGONAL PLANAR
SP ³	4	109.5°	TETRAHEDRAL

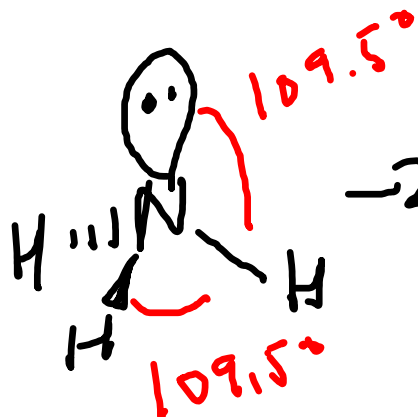


NO HYBRIDIZATION
JUST S ORBITALS

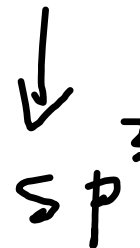


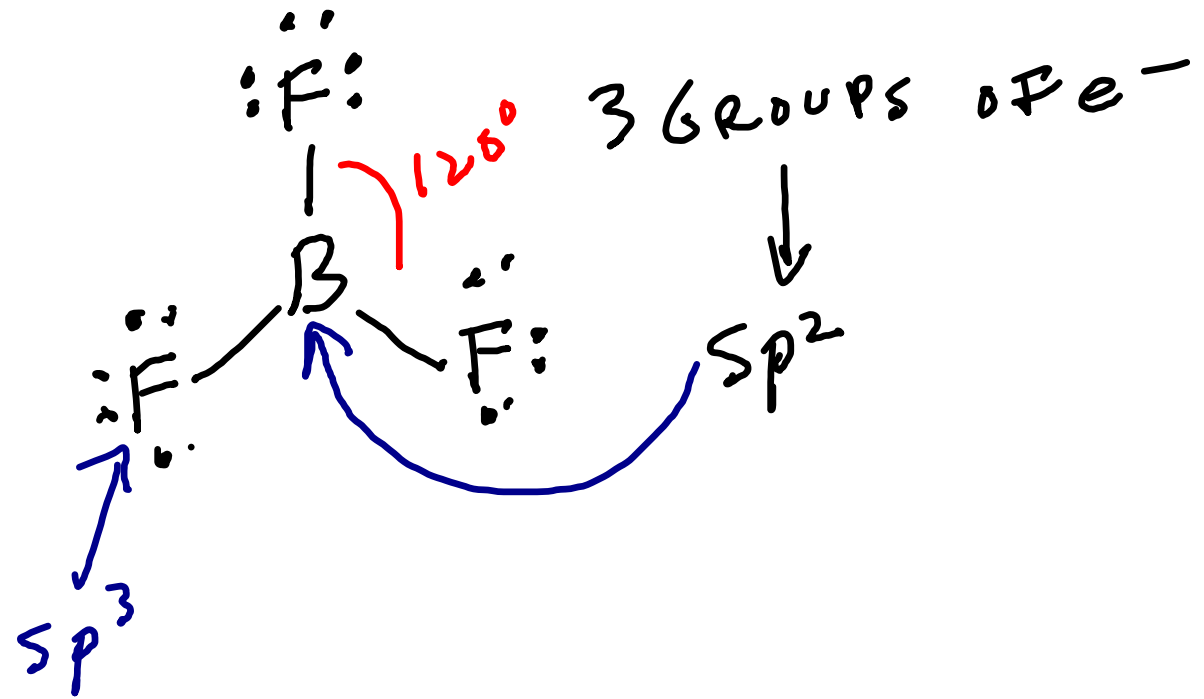
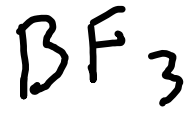
4 GROUPS of e⁻

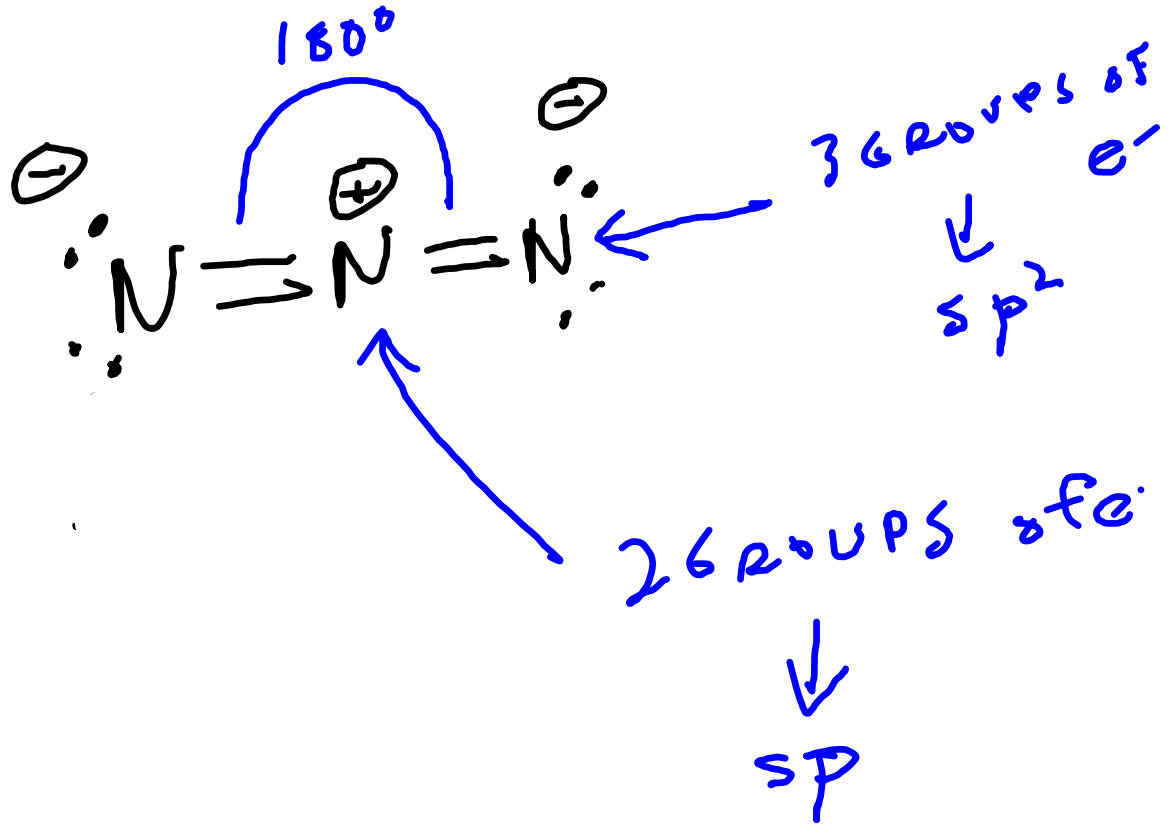


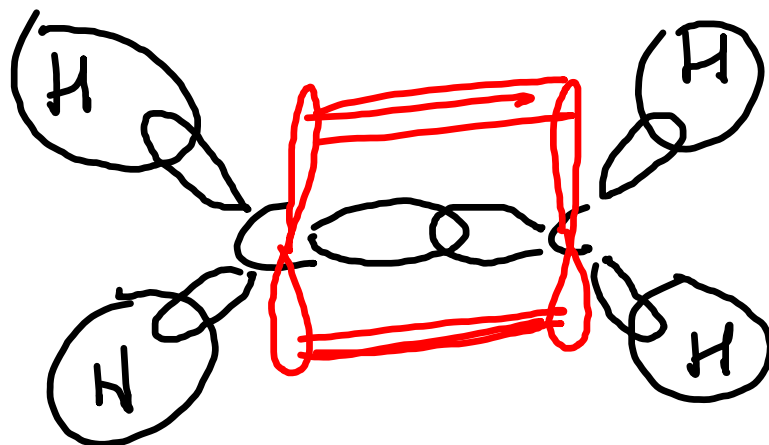
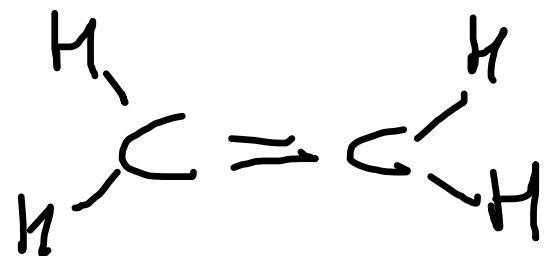


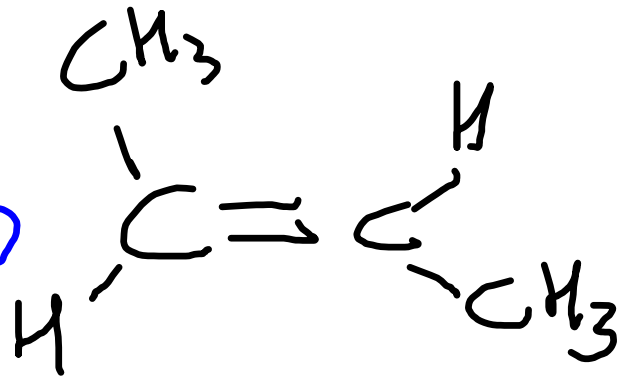
\rightarrow 4 GROUPS OF e^-



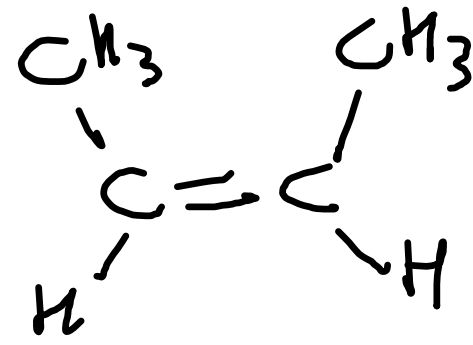








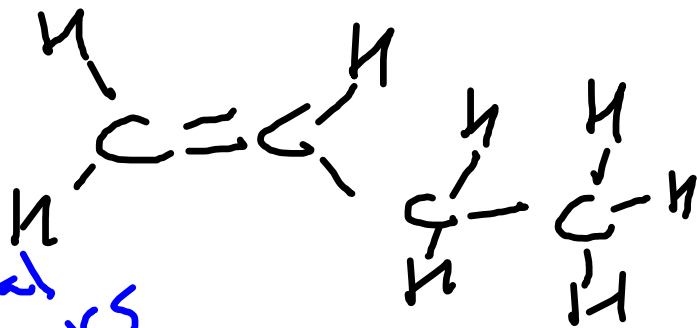
trans
2-butene



cis
2-butene

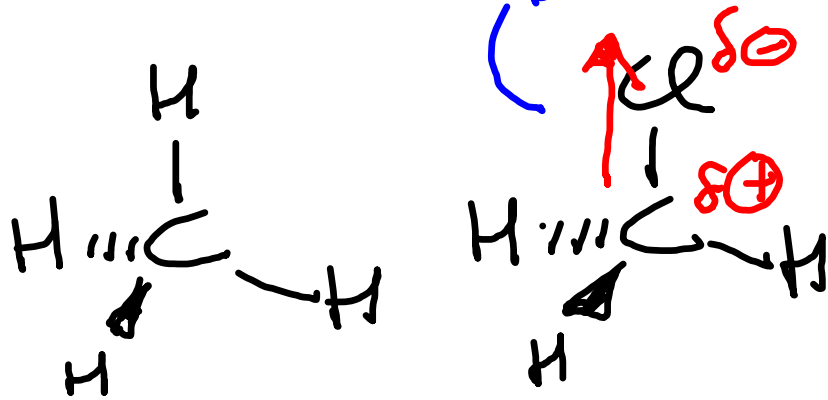
GEOMETRICAL ISOMERS

Structural
isomers



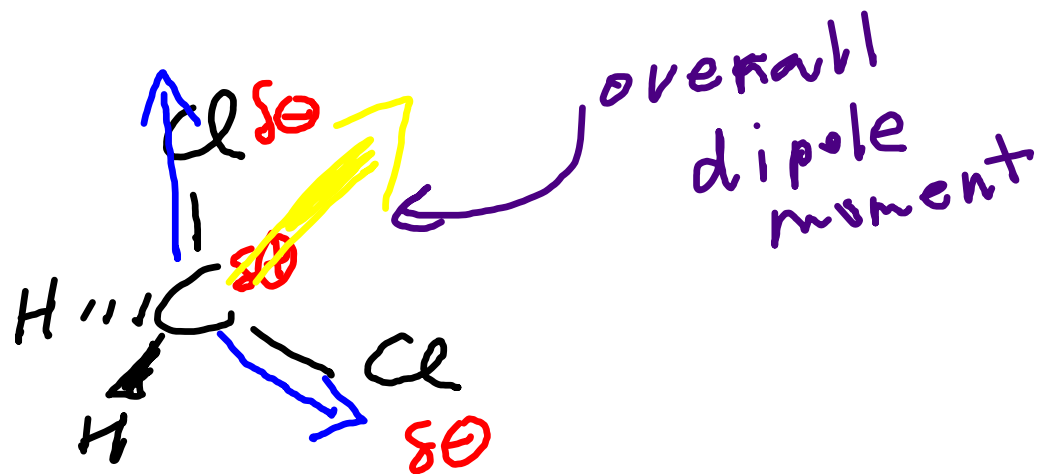
1-butene

BOND POLARITY

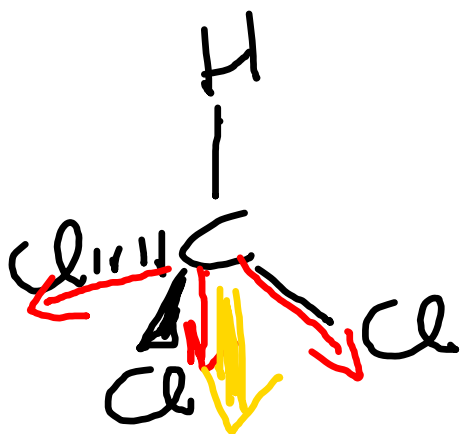


BONDS
NON POLAR
MOLECULE
NON POLAR

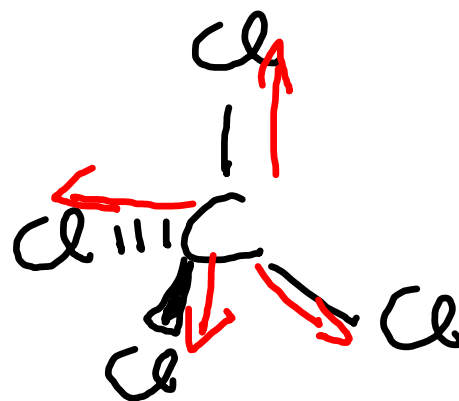
ONE
POLAR BOND
MOLECULE
POLAR



2 polar bonds
molecule
polar

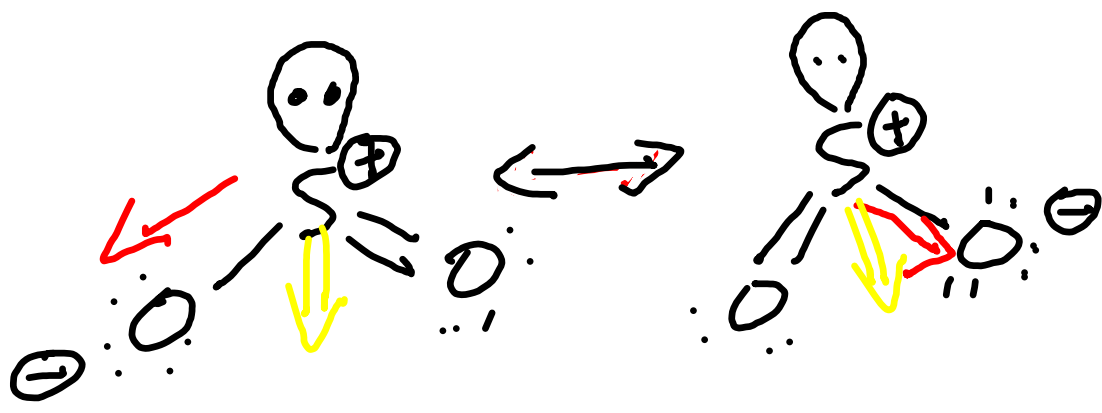


3 polar bonds
molecule
polar



4 polar bonds
molecule
non
polar

SO₂



SO_3 non polar

